

BOOKLET CODE

C

Invigilators Signature

ENTRANCE EXAMINATION – 2011
M. Sc. Chemistry

TIME: 2 HOURS

MAXIMUM MARKS: 100

HALL TICKET NUMBER:

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BOOKLET CODE:

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INSTRUCTIONS

1. Write your **HALL TICKET NUMBER** and the **BOOKLET CODE** in the space provided above and also in the **OMR ANSWER SHEET** given to you.
2. Make sure that pages numbered from 1 – 18 are present (excluding pages assigned for rough work).
3. There are 100 questions in this paper. All questions carry equal marks.
4. **There is negative marking. Each wrong answer carries - 0.33 mark**
5. Answers are to be marked on the OMR answer sheet following the instructions provided there upon.
6. Hand over both the question paper booklet and OMR answer sheet at the end of the examination.
7. In case of a tie, the marks obtained in the first 25 questions (**PART-A**) will be used to determine the order of merit.
8. No additional sheets will be provided. Rough work can be done in the space provided at the end of the booklet.
9. Only non-programmable calculators are allowed.
10. Useful constants are provided on top of **PART-A** in the question paper.

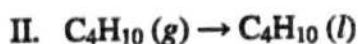
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Useful Constants:

Rydberg constant = 109737 cm^{-1} ; Faraday constant = 96500 C ; Planck constant = $6.625 \times 10^{-34} \text{ J s}$;
Speed of light = $2.998 \times 10^8 \text{ m s}^{-1}$; Boltzmann constant = $1.380 \times 10^{-23} \text{ J K}^{-1}$; Gas constant = $8.314 \text{ J K}^{-1} \text{ mol}^{-1}$;
Mass of electron = $9.109 \times 10^{-31} \text{ kg}$; Mass of proton = $1.672 \times 10^{-27} \text{ kg}$; Charge of electron = $1.6 \times 10^{-19} \text{ C}$;
 $1 \text{ D} = 3.336 \times 10^{-30} \text{ C m}$; $1 \text{ bar} = 10^5 \text{ N m}^{-2}$; $RT/F = 0.059 \text{ V}$

PART - A

1. Predict the sign of ΔS for both of the following processes



(A) ΔS should be negative for I and positive for II

(B) ΔS should be negative for I and negative for II

(C) ΔS should be positive for I and positive for II

(D) ΔS should be positive for I and negative for II

2. The remainder of $\frac{x^4 + x^3 + x^2 + x + 1}{x - 1}$ is

(A) 1

(B) 3

(C) 5

(D) 7

3. Which two colors of light cause the highest rate of photosynthesis?

(A) Red and green

(B) Red and blue

(C) Blue and green

(D) Green and yellow

4. The value of $(1)^i$ is

(A) -1

(B) i

(C) $e^{-2\pi}$

(D) $e^{-\pi}$

5. The compounds ZnO and FeO show

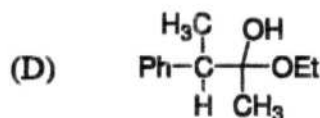
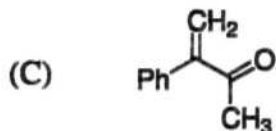
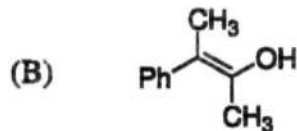
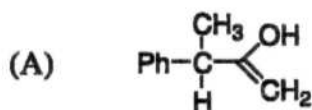
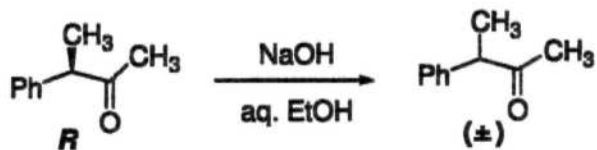
(A) metal excess and metal deficiency defects, respectively.

(B) stoichiometric and metal excess defects, respectively.

(C) metal deficiency and metal excess defects, respectively.

(D) metal excess and stoichiometric defects, respectively.

6. An intermediate in racemization of (*R*)-3-phenyl-2-butanone is



7. The packing efficiency in the hcp, bcc and simple cubic (sc) lattices are in the order

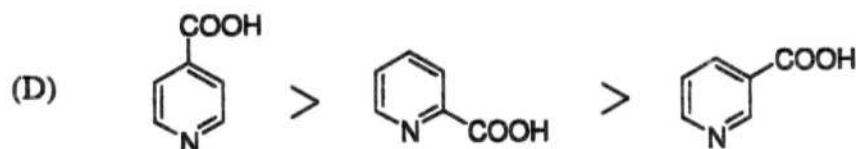
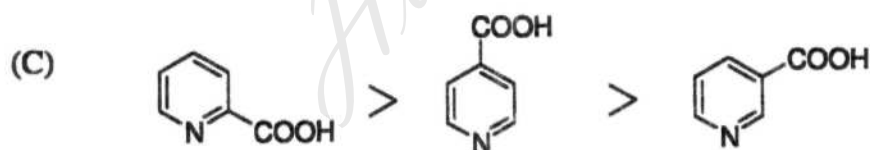
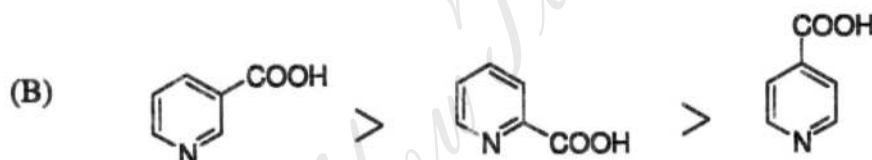
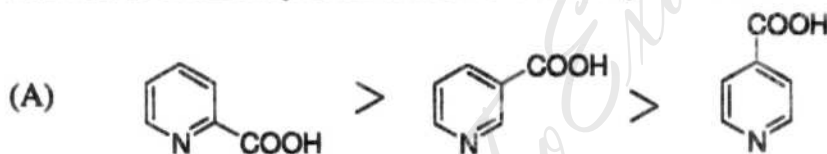
(A) bcc > hcp > sc

(B) hcp > sc > bcc

(C) hcp > bcc > sc

(D) sc > hcp > bcc

8. The rate of decarboxylation of isomeric carboxylic acids is

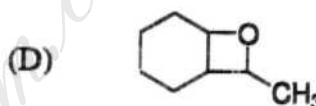
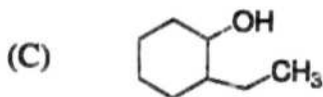
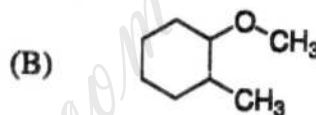
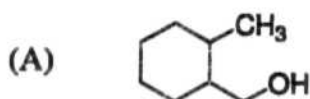
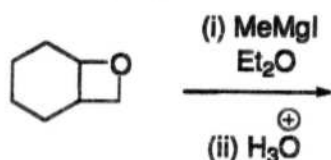


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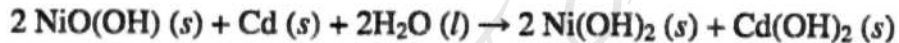
9. Doubling all the coefficients in the equation for a cell reaction

- (A) doubles ΔG^0 , but does not change E^0
- (B) doubles E^0 , but does not change ΔG^0
- (C) does not change E^0 or ΔG^0
- (D) doubles both E^0 and ΔG^0

10. The major product of the following reaction is



11. The nickel-cadmium cell has a standard potential of + 1.20 V. The cell reaction is



What is the standard free energy change for this reaction?

- (A) - 116 kJ
- (B) - 38.7 kJ
- (C) - 232 kJ
- (D) - 46.3 kJ

12. If $\cos A = x$; then $\cos 4A =$

- (A) $4x$
- (B) $8x^4 - 8x^2 + 1$
- (C) $4x^4 - 4x^2 + 1$
- (D) $2x^2 + 1$

13. Which of the following is necessary for a process to be spontaneous ($\Delta S =$ change in entropy)?

- (A) $\Delta S_{\text{universe}} > 0$
- (B) $\Delta S_{\text{system}} < 0$
- (C) $\Delta S_{\text{system}} > 0$
- (D) $\Delta S_{\text{surroundings}} < 0$

14. If two vertices of a cube chosen randomly are painted black and the remaining are painted white, what is the probability that the black vertices are adjacent i.e. connected by an edge?

- (A) $\frac{2}{7}$ (B) $\frac{3}{7}$ (C) $\frac{1}{2}$ (D) $\frac{3}{28}$

15. CuI_2 is unstable and it readily decomposes to

- (A) Cu and I^- (B) Cu and I_2 (C) CuI and I^- (D) CuI and I_2

16. Consider the plot of the function $y = 1/x$. The tangent to this curve drawn at the point (1, 1), will cut the x-axis at:

- (A) (2, 0) (B) ($\sqrt{2}$, 0) (C) (1, $\sqrt{2}$) (D) (1, 0)

17. A sample of water contains 200 ppm of Ca^{2+} . What is the molality of the solution with respect to Ca^{2+} ? Atomic weight of Ca is 40.

- (A) 0.2 m (B) 2 m (C) 0.05 m (D) 5×10^{-3} m

18. The strongest Brønsted acid among the following is

- (A)  (B) 
- (C)  (D) 

19. Which one among the following chlorides is dissociated to the least extent in aqueous solution?

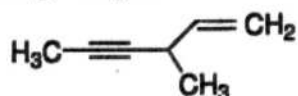
- (A) HgCl_2 (B) ZnCl_2 (C) BaCl_2 (D) AlCl_3

20. Which of the following is not a crystalline substance?

- (A) Diamond (B) Graphite (C) Charcoal (D) C_{60}

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21. The IUPAC name for the following compound is

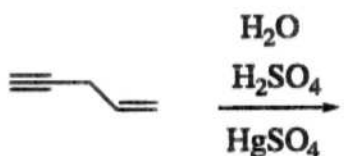


- (A) 4-vinyl-2-pentyne (B) 4-methylhex-2-yn-5-ene
 (C) 3-methylhex-1-en-4-yne (D) 3-methylhex-4-yn-1-ene

22. X-ray diffraction study of a crystal with a simple cubic lattice structure shows diffraction from the (110) plane appearing at the Bragg angle $\theta = 20^\circ$. The angle at which the diffraction from the (220) plane will appear is

- (A) 9.8° (B) 10° (C) 40° (D) 43°

23. The major product expected from the following reaction is



- (A) (B)
 (C) (D)

24. Consider the equilibrium $X \rightleftharpoons 2Y$ with equilibrium constant, $K_C = 3.6$ M at 25°C . If the initial concentrations are $[X]_0 = 1.0$ M and $[Y]_0 = 0.0$ M, the equilibrium concentration of X at 25°C , $[X]_{\text{eq}}$ is

- (A) 0.33 M (B) 0.40 M (C) 0.36 M (D) 0.60 M

25. The sides of a triangle are of length 3.0, 4.0 and 5.0 cm. If the side with length 5.0 cm is the base, what is the height of the triangle?

- (A) 2.4 cm (B) 2.8 cm (C) 3.4 cm (D) 4.0 cm

PART - B

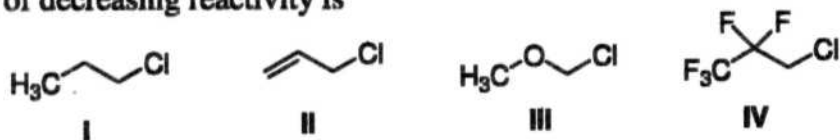
26. The heat of reaction of both the reactions $2 \text{KOH} + \text{H}_2\text{SO}_4 \rightarrow \text{K}_2\text{SO}_4 + 2\text{H}_2\text{O}$ and $\text{Mg}(\text{OH})_2 + \text{H}_2\text{SO}_4 \rightarrow \text{MgSO}_4 + 2 \text{H}_2\text{O}$ is -27.2 kcal/mol . Hence the heat of reaction of $3 \text{Ca}(\text{OH})_2 + 2 \text{H}_3\text{PO}_4 \rightarrow \text{Ca}_3(\text{PO}_4)_2 + 6 \text{H}_2\text{O}$ would be
- (A) -13.6 kcal/mol (B) -27.2 kcal/mol
 (C) -68.0 kcal/mol (D) -81.6 kcal/mol
27. The standard equation of a circle passing through the points $u (3, 8)$, $v (9, 6)$ and $w (13, -2)$ is
- (A) $(x - 2)^2 + (y + 3)^2 = 100$ (B) $(x + 3)^2 + (y + 2)^2 = 100$
 (C) $(x - 3)^2 + (y + 2)^2 = 100$ (D) $(x - 2)^2 + (y - 3)^2 = 100$
28. Acid is used in the standardization titration of KMnO_4 against sodium oxalate because
- (A) it helps in dissolving KMnO_4 .
 (B) it stabilizes permanganate ion.
 (C) it facilitates the reduction of Mn^{7+} to Mn^{2+} .
 (D) it helps in dissolving the MnO_2 formed during titration.
29. Which of the following covalent compounds does not have any formally charged atom?
- (A) $(\text{CH}_3)_3\text{NO}$ (B) CH_2N_2 (C) CH_3ONO (D) CH_3CNO
30. The two radial nodes in the $3s$ radial function of H atom occur at the distances R_1 and R_2 from the nucleus. The three radial nodes in the $4s$ orbital occur at R_3 , R_4 and R_5 . The order of these distances is given by
- (A) $R_3 < R_1 < R_4 < R_2 < R_5$ (B) $R_1 < R_3 < R_4 < R_2 < R_5$
 (C) $R_3 < R_1 < R_2 < R_4 < R_5$ (D) $R_3 < R_1 < R_4 < R_5 < R_2$
31. The graph of the equation $4(x^2 - 4x) - 9(y^2 - 2y) - 29 = 0$ represents a
- (A) parabola (B) ellipse (C) circle (D) hyperbola

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32. Using Wade's rule predict the structure of B_5H_9

- (A) closo (B) arachno (C) nido (D) scorpionato

33. In the S_N1 solvolysis of the following primary alkyl chlorides in aqueous ethanol, the order of decreasing reactivity is



- (A) I > II > III > IV (B) II > I > III > IV
(C) IV > III > II > I (D) III > II > I > IV

34. A solution of sulfuric acid contains 86 g of H_2SO_4 per liter of solution. The normality of the solution is

- (A) 0.9 N (B) 1.8 N (C) 2.0 N (D) 1.0 N

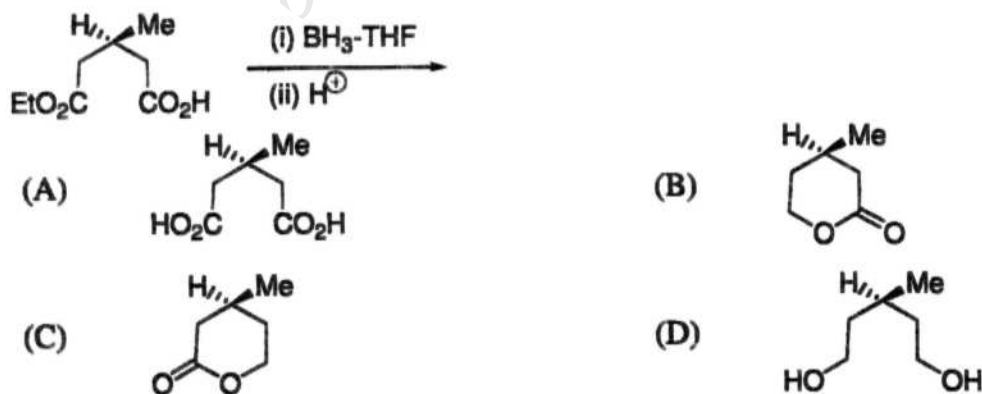
35. The equation of the normal line to $y = x^3 - 2x^2 + 4$ at (2, 4) is

- (A) $y = -\frac{1}{4}x + \frac{9}{2}$ (B) $y = 9x + 4$
(C) $y = -4x + \frac{9}{2}$ (D) $y = -9x + \frac{1}{4}$

36. An extensive property of a thermodynamic system among the following is

- (A) Pressure (B) Temperature (C) Volume (D) Concentration

37. The product obtained in the following transformation is



38. The number of unpaired electrons in the complex ion is in the order

- (A) $\text{ZnCl}_4^{2-} < \text{CuCl}_4^{2-} < \text{NiCl}_4^{2-} < \text{CoCl}_4^{2-}$
 (B) $\text{ZnCl}_4^{2-} < \text{CuCl}_4^{2-} < \text{CoCl}_4^{2-} < \text{NiCl}_4^{2-}$
 (C) $\text{ZnCl}_4^{2-} < \text{CoCl}_4^{2-} < \text{NiCl}_4^{2-} < \text{CuCl}_4^{2-}$
 (D) $\text{CuCl}_4^{2-} < \text{NiCl}_4^{2-} < \text{ZnCl}_4^{2-} < \text{CoCl}_4^{2-}$

39. If $F(x) = x^{1/x}$ then $\lim_{x \rightarrow \infty} F(x) =$

- (A) 0 (B) 1 (C) ∞ (D) e

40. The metal ion involved in the water oxidation process at the active site of photosystem II is

- (A) Fe (B) Mg (C) Mo (D) Mn

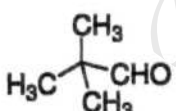
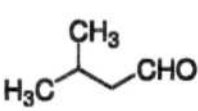
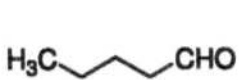
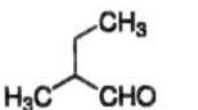
41. The oxidation number of carbon in dimethyl ether is

- (A) -1 (B) -2 (C) 1 (D) 2

42. The complex number $-2 - 2\sqrt{3}i$ in polar form is given by

- (A) $2e^{i2\pi/3}$ (B) $4e^{i2\pi/3}$ (C) $4e^{i3\pi/2}$ (D) $4e^{i4\pi/3}$

43. Compound I gives a strong infrared absorption at 1730 cm^{-1} . ^1H NMR spectrum indicates that it has two types of hydrogen atoms; one H atom appearing as singlet at $\delta = 9.7 \text{ ppm}$ and 9 H atoms appearing as a singlet at $\delta = 1.2 \text{ ppm}$. The structure of I is

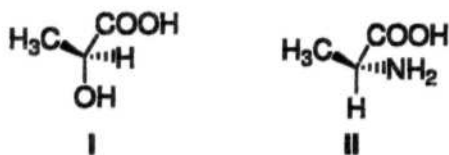
- (A)  (B) 
- (C)  (D) 

44. In a titration experiment, the end point indicates

- (A) apparent equivalence point (B) completion of precipitation
 (C) neutralization point (D) exact equivalence point

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45. The structures I and II, shown below, correspond to:



- (A) *S*-lactic acid and *S*-alanine (B) *R*-lactic acid and *R*-alanine
 (C) *R*-lactic acid and *S*-alanine (D) *S*-lactic acid and *R*-alanine

46. The magnetic quantum number of the last electron in the atom with atomic number 21 is

- (A) 4 (B) 2 (C) 3 (D) 1

47. If the number e^{9i} is marked as a point on the complex plane, what is the distance of the point from the origin?

- (A) 1 (B) 3 (C) 9 (D) $\tan^{-1} 9$

48. What is the electronic configuration of V^{3+} ?

- (A) $[Ar]3d^2$ (B) $[Kr]3d^2$ (C) $[Ar]3d^3$ (D) $[Kr] 3d^3$

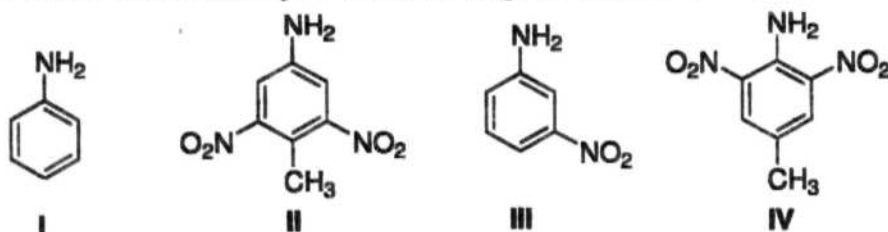
49. The hybridization that is common for at least one of the carbon atoms in hydrogen cyanide, carbon disulfide, allene and carbon monoxide is

- (A) sp (B) sp^2 (C) sp^3 (D) dsp^3

50. Which one among the given functions has the smallest slope at $x = 1$?

- (A) $2x^2 - 3$ (B) $2x^2 - 1$ (C) $2x^2 - 2x$ (D) $2x^2 - x$

51. The order of basicity of the following substituted anilines is



- (A) I > II > III > IV (B) IV > I > III > II
 (C) IV > II > I > III (D) IV > III > I > II

52. According to crystal field theory, the 9th electron of the metal centre in square planar $[\text{Cu}(\text{NH}_3)_4]^{2+}$ resides in
 (A) d_{xy} (B) d_{z^2} (C) $d_{x^2-y^2}$ (D) d_{xz}
53. The rotation of pure *R*(+)-Limonene is $+123.0^\circ$. Calculate the % of the (+) isomer in a sample showing a rotation of $+109^\circ$?
 (A) 8.6 (B) 77.2 (C) 94.3 (D) 82.9
54. An ideal gas undergoes isothermal and reversible expansion from its initial volume to some final volume at 300 K drawing in 90 kJ of heat. The change in the Gibbs free energy of the gas is
 (A) 300 J (B) 0 J (C) 150 J (D) -300 J
55. For $0 \leq x \leq 1$, $\lim_{n \rightarrow \infty} \int_0^1 n x e^{-nx^2} dx =$
 (A) $\frac{1}{2}$ (B) ∞ (C) 0 (D) 2
56. The two strands of double helical DNA are associated by hydrogen bonds between adenine (A) and thymine (T), and between guanine (G) and cytosine (C). The numbers of hydrogen bonds between A-T and G-C pairs, respectively are:
 (A) one, two (B) two, two (C) three, two (D) two, three
57. The intermediate involved in Curtius rearrangement is
 (A) carbenium ion (B) carbanion
 (C) nitrene (D) carbene
58. An organic compound on decomposition at 500°C and 1 atm. pressure released 2 mL each of carbon monoxide, nitrogen and water vapour. The empirical formula of the molecule is
 (A) CHNO (B) $\text{CH}_2\text{N}_2\text{O}_2$ (C) $\text{CH}_2\text{N}_2\text{O}$ (D) CH_2NO

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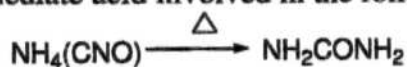
59. $\frac{d}{dt} \exp[t^2] =$

- (A) $1/t^2$ (B) $\exp(t^2)$
 (C) $2t \exp(t^2)$ (D) $(1/2t) \exp(t^2)$

60. Nessler's reagent is prepared by mixing a solution of KI with a solution of 'X' and then adding KOH solution. Here 'X' is

- (A) $ZnCl_2$ (B) $HgCl_2$ (C) $AlCl_3$ (D) $TiCl_4$

61. The intermediate acid involved in the following reaction is



- (A) cyanic acid (B) cyanuric acid (C) thiocyanic acid (D) uric acid

62. 50 mL of 0.04 M HCl solution was mixed with 50 mL of 0.02 M $AgNO_3$ solution, stirred and filtered. The pH of the filtered solution is

- (A) 1 (B) 4 (C) 3 (D) 2

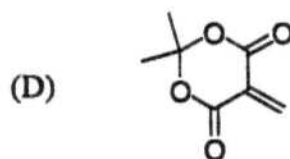
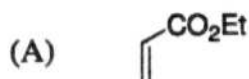
63. $(1 + 2i)^{-1}$ is equal to

- (A) $1 - 2i$ (B) $(1/5) - (2/5)i$
 (C) $(1/3) - (2/3)i$ (D) $-(1/3) + (2/3)i$

64. Which of the following fluorides is angular?

- (A) SnF_2 (B) ZnF_2 (C) BeF_2 (D) XeF_2

65. The more reactive dienophile among the following for the reaction with cyclopentadiene is



66. Given that ^{18}F undergoes 90 % radioactive decay in 366 min., the half life ($t_{1/2}$) for ^{18}F is
 (A) 110 min. (B) 3473 min. (C) 154 min. (D) 220 min.
67. The phenolic compound among the following is:
 (A) Ibuprofen (B) Camphor (C) Penicillin (D) Paracetamol
68. What is the hydroxide ion concentration of a solution that has a pH of 11.20?
 (A) 6.31×10^{-12} M (B) 11.20 M (C) 1.58×10^{-3} M (D) 2.80 M
69. For all values of x which determinant among the following is zero?
 (A) $\begin{vmatrix} x & 1 \\ 1 & x \end{vmatrix}$ (B) $\begin{vmatrix} 1 & x \\ x & x^2 \end{vmatrix}$ (C) $\begin{vmatrix} 1 & x \\ x & 1 \end{vmatrix}$ (D) $\begin{vmatrix} 1 & x^2 \\ x & 1 \end{vmatrix}$
70. The conductivity of sodium dodecyl sulfate (SDS) solution exhibits a sharp transition around 8 mM concentration. This is because:
 (A) SDS precipitates beyond 8 mM concentration.
 (B) SDS forms micelles above 8 mM concentration.
 (C) SDS forms a gel above 8 mM concentration.
 (D) SDS undergoes hydrolysis above 8 mM concentration.
71. According to the equation $2\text{Fe}^{3+} + 2\text{I}^- \longrightarrow \text{I}_2 + 2\text{Fe}^{2+}$
 how many grams of iodine can be produced by reacting 7.4 mols of Fe^{3+} and 7.0 mols of I^- ? [At. Wt. of iodine is 127]
 (A) 9.1×10^2 g (B) 8.9×10^2 g (C) 9.4×10^2 g (D) 17.8×10^2 g
72. The most appropriate spectroscopy for the identification of a nitrile group is
 (A) IR (B) ^1H NMR (C) UV (D) ESR
73. If the units for rate are M s^{-1} , what are the units for the rate constant, k , for a zeroth-order reaction?
 (A) M s^{-1} (B) $\text{M}^{-1}\text{s}^{-1}$ (C) s^{-1} (D) M^{-1}

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74. The function with a finite range is
(A) e^x (B) e^{x^2} (C) e^{x^3} (D) e^{-x^2}
75. How many grams of copper will be produced when 27 g of aluminium is added to excess cupric sulphate solution? [At. wts.; Al = 27, Cu = 63.5]
(A) 63.50 (B) 95.25 (C) 90.50 (D) 122.25
76. Assuming the additivity of covalent radii [C 0.77 Å, Br 1.14 Å], and assuming the distance between adjacent carbon atoms in the ring as 1.40 Å, the distance between the centres of bromine atoms in 1,2-dibromobenzene is
(A) 3.31 Å (B) 3.42 Å (C) 4.20 Å (D) 2.28 Å
77. The number of stereoisomers for CHD=CH-CH=CHD is
(A) 6 (B) 8 (C) 2 (D) 4
78. The entropy change associated with the expansion of one mole of an ideal gas from an initial volume of V to a final volume of $2.50 V$ at constant temperature is (R = gas constant)
(A) $\Delta S = -R \ln 2.50$ (B) $\Delta S = R \ln 2.50$
(C) $\Delta S = 2.50 R \ln (V_f/V_i)$ (D) $\Delta S = -2.50 R \ln (V_f/V_i)$
79. The smallest among the following integrals is
(A) $\int_0^1 e^{-x^4} dx$ (B) $\int_0^1 e^{-x^2} dx$ (C) $\int_0^1 e^{-x^3} dx$ (D) $\int_0^1 e^{-x} dx$
80. The quaternary structure of human hemoglobin is a
(A) dimer of two myoglobin dimers.
(B) tetramer of identical subunits.
(C) tetramer of four different subunits.
(D) tetramer of two different subunits.

81. The number of isomers having non-zero dipole moment for PCl_2F_3 in the trigonal bipyramidal geometry is
(A) 3 (B) 2 (C) 1 (D) 0
82. The most appropriate reagent for the conversion of RCOOMe into RCH_2OH is
(A) NaBH_4 (B) NaH (C) LiBH_4 (D) Pd/C and H_2
83. Which of the following statements must be true for the entropy of a pure solid to be zero?
(I) The temperature must be 0 K.
(II) The solid must be crystalline, not amorphous.
(III) The solid must be perfectly ordered.
(IV) The solid must be an element.
(A) I, II and III (B) I and II (C) I (D) I, II, III and IV
84. The function with exactly two minima and one maximum, among the following is
(A) $x + x^2 - x^4$ (B) $x^4 - x^2 - x$ (C) $x^3 - x^2 - x$ (D) $x + x^2 - x^3$
85. Collagen is
(A) an α -helical structural protein.
(B) a coiled-coil protein found in hair.
(C) a cross-linked globular protein.
(D) a triple-helical fibrous protein.
86. How many milliliters of 2 M NaCl solution are required to make one litre of 0.4 M NaCl solution by adding water?
(A) 5000 ml (B) 800 ml (C) 200 ml (D) 20 ml

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87. Which of the following compounds is aromatic?



88. A 0.01 M solution of a compound transmits 20 % of visible light when the absorbing path length is 1.5 cm. What is the molar extinction co-efficient of the substance? Solvent is assumed to be completely transparent.

- (A) $46.6 \text{ M}^{-1}\text{cm}^{-1}$ (B) $50.3 \text{ M}^{-1}\text{cm}^{-1}$ (C) $22.3 \text{ M}^{-1}\text{cm}^{-1}$ (D) $43.6 \text{ M}^{-1}\text{cm}^{-1}$

89. Which of the following atoms has the highest number of unpaired electrons in its ground state?

- (A) N (B) C (C) O (D) F

90. Which of the following compounds has the highest boiling point?

- (A) Toluene (B) Benzene
(C) Mesitylene (D) Cyclohexane

91. Which one of the following statements do not apply to interhalogen compounds?

- (A) Could be neutral (B) Could be cationic
(C) Could be anionic (D) Always obey octet rule.

92. The product obtained by the reaction of one equivalent of 1-bromo-3-chlorocyclobutane and two equivalents of Na is:



93. Which of the following pair has the lowest interfacial tension?

- (A) *n*-decane/water (B) *n*-octyl alcohol/water
 (C) air/water (D) *n*-butane/water

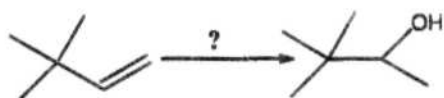
94. The gas pressure in an aerosol container is 1.5 atm at 25°C. Assuming an ideal behavior of the gas, if the container is heated to 450°C, the pressure would be close to

- (A) 1.023 atm (B) 1.234 atm (C) 3.639 atm (D) 2.639 atm

95. The order of increasing dipole moment among H₂S, H₂O and BF₃ is

- (A) BF₃ < H₂S < H₂O (B) H₂O < H₂S < BF₃
 (C) H₂S < H₂O < BF₃ (D) BF₃ < H₂O < H₂S

96. The best method for the following transformation is



- (A) acid mediated hydration (B) hydroboration-oxidation
 (C) ozonolysis-reduction (D) oxymercuration-demercuration

97. The concentration of Ba²⁺ in saturated BaSO₄ solution at 27°C is 1.04 × 10⁻⁵ M. What is the solubility product (K_{sp}) for BaSO₄ at this temperature?

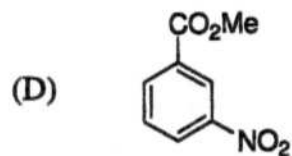
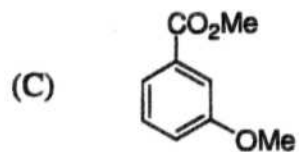
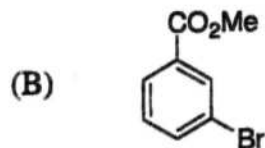
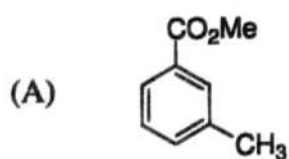
- (A) 1.04 × 10⁻¹⁰ M (B) 2.08 × 10⁻⁵ M
 (C) 0.52 × 10⁻¹⁰ M (D) 1.08 × 10⁻¹⁰ M

98. What is the hybridization of sulfur in SF₄?

- (A) sp² (B) sp³d (C) sp³ (D) sp²d²

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99. The ester that undergoes acid hydrolysis most readily is



100. If the half-life of a reaction is independent of its initial concentration, then the reaction may be categorized as

(A) zeroth order

(B) first order

(C) second order

(D) bimolecular