



SB-1154
Second Year B. Pharm. Examination
March / April – 2011
PH - 204 : Pharmaceutical Analysis - I

Time : 3 Hours]

[Total Marks : 70

Instructions :

(1)

<p>नीचे दृशावेल निशानीवाणी विगतो उत्तरवडी पर अवश्य लपवी. Fillup strictly the details of signs on your answer book.</p> <p>Name of the Examination :</p> <p>Second Year B. Pharm.</p> <p>Name of the Subject :</p> <p>PH - 204 : Pharmaceutical Analysis - 1</p> <p>Subject Code No. : 1 1 5 4 Section No. (1, 2,.....) : 1&2</p>	<p>Seat No. :</p> <table border="1" style="width: 100%; height: 20px;"><tr><td style="width: 15%;"></td><td style="width: 15%;"></td><td style="width: 15%;"></td><td style="width: 15%;"></td><td style="width: 15%;"></td><td style="width: 15%;"></td></tr></table> <div style="border: 1px solid black; border-radius: 15px; padding: 10px; margin-top: 10px; text-align: center;">Student's Signature</div>						

SECTION - I

- 1 Attempt any five : 10**
- (1) Define and classify volumetric analysis.
 - (2) Explain Lowry-Bronsted concept of Acid-Base with examples.
 - (3) Classify the solvents used in non-aqueous titrations with examples.
 - (4) Explain Warner's coordination numbers.
 - (5) Comment on the statement "Aqueous solution of Sodium acetate is neutral".
 - (6) Calculate pH and pOH of a $5.0 \times 10^{-2} M$ solution of NaOH.
- 2 Attempt any three : 9**
- (1) Explain various types of complexometric titrations.
 - (2) Explain various factors affecting solubility product constant.
 - (3) What is co-precipitation ? How will you minimize co-precipitation ?
 - (4) Explain Volhard's method of precipitation titration.
- 3 Attempt any four : 16**
- (1) Explain leveling and differentiating effects of solvents.
 - (2) Discuss permanganate titration in acidic and basic condition using suitable example.

- (3) Define and Classify errors in pharmaceutical analysis.
- (4) Write a brief note on acid base indicators.
- (5) Write a note on adsorption indicator.

SECTION - II

- 4 (a) Explain following terms with examples : (any **four**) **8**
- (1) Iodometry
 - (2) Buffer capacity
 - (3) Solubility product constant
 - (4) Primary standard compound
 - (5) Common ion effect
 - (6) Post precipitation.
- (b) Explain accuracy and precision with suitable examples. **3**
- 5 Attempt any **three** : **12**
- (1) Application of masking agent.
 - (2) Discuss the factors affecting solubility of precipitate.
 - (3) Distinguish between back titration and blank titration.
 - (4) Define buffers and give its applications in pharmacy.
- 6 Attempt any **two** : **12**
- (1) Write a note on Oxygen flask combustion method.
 - (2) Write a note on diazotization titration.
 - (3) Discuss Kjeldahl method of nitrogen estimation in detail.
-